

Pool chemistry: Backyard perils and experiments

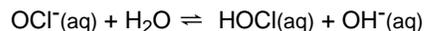
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One of the first things I noticed when I bought a pool is that pool chemistry terms are confusing to both lay people and specialists alike. A pool manual's usage of terms such as "alkalinity" and "free chlorine" are not textbook definitions. But once I sorted through the initial confusion, it gave me one more excuse to do some simple experiments in my backyard.

1. Chlorine and pH

When pool manufacturers use the term "free chlorine", they are referring to hypochlorous acid, HOCl, which exists in aqueous solution. This is the chemical species that actually kills bacteria and attenuates algal growth, preventing diarrhea, swimmer's ear (a nasty earache) and various skin and respiratory problems. Diatomic chlorine, Cl₂, would also do the trick, but it is far too dangerous.

If Ca(ClO)₂ is used as an alternative, the product will slowly dissolve in water to produce the hypochlorite ion, OCl⁻(aq). But OCl⁻(aq) goes into equilibrium with the vital HOCl.



If the pH is too high, the reverse reaction will be encouraged, leading to a higher concentration of OCl⁻ and lowering the amount of essential HOCl. Although a lower pH (more acid) will encourage the forward reaction (by destroying OH⁻), it also makes the water more irritating and more corrosive. Copper for example, which is used in

some electric heating coils and in less expensive heat pumps, can withstand the action of HOCl alone. Unfortunately the metal is prone to oxidation in the presence of both HOCl and acid. The right balance of effective bactericidal action and low corrosion occurs at a pH of 7.2.

The ideal concentration of hypochlorite in pool water is about 2 parts per million, translating into about 2 mg of calcium hypochlorite per liter of water. [See page 10 for an activity on parts per million]. I normally start with a little more, especially if I'm not chlorinating at dusk (best time). Ultraviolet light will break down chlorine compounds, create Cl free radicals and eventually convert them to HCl. If I dissolve 120 g in a 40,000-L pool, that's equivalent to 3 g per 1000 L, which is 0.003 g/L or 3 mg per L.

On average, such an amount will last two days. But factors such as temperature, frequency of use, and presence of a pool cover (which filters UV light) will either increase or decrease HOCl's residence-time. There are times in the summer when my HOCl runs low after 24 hours. Conversely, in the spring before we actually start swimming in the pool, I have seen chlorine levels remain adequate beyond 48 hours.

Adding more than the required amount of hypochlorite is not recommended because it is an eye and skin irritant.

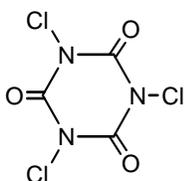
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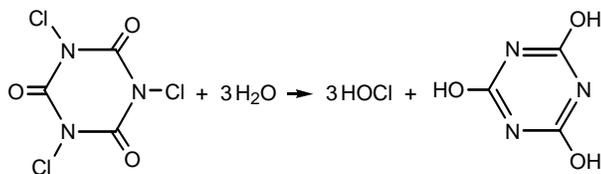
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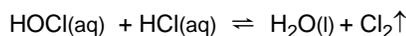
For those seeking a more stable chlorine-product and the illusion of convenience, *trichlor* (trichloroisocyanuric acid) tablets (or “pucks”) are available. In addition to generating HOCl, trichloroisocyanuric acid lowers pH and drives the HOCl-equilibrium to the right (good old LeChatelier principle), but unless the owner periodically adds carbonate ion (so-called pH⁺), the acid will accumulate and take its toll on pool equipment. It is for this reason that I prefer Ca(ClO)₂. Although it too has its drawbacks (calcium deposits can accumulate if dilution and backwashes are not routinely performed) it is nowhere near as nasty as trichlor.



Thanks to trichlor and some carelessness on my part, a couple of years ago I almost became a World War I-like victim of chlorine gas. Here's how. First trichlor creates HOCl and isocyanuric acid:



With both acid and HOCl around, elemental chlorine can potentially be part of the following equilibrium:



Normally the acid introduced by the pucks, once diluted by a pool's large volume of water, will not create much of the elemental chlorine gas. But, I had a leak in my storage shed, and somehow water crept past what I thought was a waterproof lid. The water in the bucket of trichloroisocyanuric acid pucks partially dissolved the chemical, creating very high concentrations of both HOCl and HCl. So when I opened the container I was greeted by the second-most vicious member of the halogen family. The greenish gas felt like a knife going through my throat and lungs. Luckily I was alone and outdoors. Good

ventilation and a strong coughing reflex flushed most of it out of my system.

2. Alkalinity

Alkalinity to a chemist means level of base concentration; a base is an acid's counterpart. But to a pool person, alkalinity refers to a specific base, namely hydrogen carbonate ion (HCO₃⁻). The reason it's important to have a decent amount of this in your pool, is that it acts as a buffer, thus stabilizing the pH level.

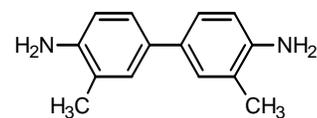
3. Measuring pH

The typical pH paper provided in some kits is not at all accurate. It can easily be off by more than 1 unit, and a single unit on the pH scale, a logarithmic one, is really a factor of 10 in terms of [H⁺]. Some kits come with phenol red, but it is useful only between pH values of 7 and 8. A pH meter would be nice, or you can embark upon the following adventure.

If the water is more acidic than 7 — tap water and rainwater typically are — you will first need bromothymol blue. Take a small sample of tap water and one of pool water. Since the indicator bromothymol blue is dissolved in NaOH, adding four drops to a small vial containing an unbuffered sample will not give the expected green colour for tap water. It will be blue. If the pool water plus four drops of the same indicator is as blue as the tap water sample then you will not need to add the carbonate at this point, go on to perform the phenol red test. If the pool sample is a lighter blue, or greenish blue, it will obviously be more acidic than tap water. This will correspond to a pH of <6.8 in the chart in the back of the pH⁺ bottle. It's time to add carbonate in proportion to the volume of the pool. Then after 12 hours, remeasure the pH using phenol red indicator. If it's still below the desired pH value of 7.2, add more carbonate. If you accidentally overshoot the desired value, it will be time to buy pH⁻, and you're on your way to having your own lab.

4. Chlorine Indicator

A common agent used to measure free chlorine is ortho-tolidine, which turns either a light or dark shade of



yellow depending on the concentration of HOCl. Ortho-tolidine is, unfortunately, a cancer-suspect agent, so I wear gloves, and I do not dump the samples back into the pool. Interestingly, the U.S. Environmental Protection Agency no longer approves it for chemical analysis, but it is still widely sold in Canada.

Since it's vital to constantly monitor pH and chlorine levels in one's swimming pool — too much chlorine is toxic; too little leads to problems; too little acid will not let HOCl do its job, and too much will be corrosive — a chlorine indicator is indispensable. Hopefully, a safer one will soon be commercially available. Otherwise, I will have to rely on the old iodide-thiosulfate titration.

References

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2. Terry Tamminen, *The Ultimate Pool Maintenance Manual*, McGraw-Hill, 2000.
3. <http://www.rhtubs.com/chlorine.htm#trichlor>
4. http://en.wikipedia.org/wiki/Calcium_hypochlorite
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A microbiologist's favourite dish

A **Petri dish** is a shallow glass or plastic cylindrical dish that biologists use to culture cells, which can be bacterial, animal, plant, or fungus. It was named after the German bacteriologist Julius Richard Petri (1852 –1921) who invented it in 1877 when working as an assistant to Robert Koch.

(From http://en.wikipedia.org/wiki/Petri_dish)

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