

Name: _____ Period: _____

Chemical Reactions and Solubility Rules

Soluble vs Insoluble Demonstration:

DRAW – what you see happening			EXPLAIN – use the words: dissolve, stir, water		
CaCl ₂	CaCO ₃	Na ₂ CO ₃	CaCl ₂	CaCO ₃	Na ₂ CO ₃

Definitions:

Soluble:

Insoluble:

Precipitate:

Solubility Rules: (Handout) Practice applying Solubility Rules

CaCl ₂	CaCO ₃	Na ₂ CO ₃
1. Identify the ions: 2. Determine which rule applies	1. Identify the ions: 2. Determine which rule applies	1. Identify the ion: 2. Determine which rule applies:

What happens when CaCl_2 and Na_2CO_3 combine?

PREDICT	OBSERVE	EXPLAIN

Lab: Remember – Safety Goggles! Only use 3-4 drops of each solution.

Reactants (3-4 drops)	Prediction of Products and solubility (using solubility chart)	Observation of Reaction		Chemical Equation
		Remains clear = no reaction	Becomes cloudy = precipitate formed	
1.				
2.				
3.				
4.				

Explain: Using the information from today, choose one chemical reaction with a precipitate and **explain why the precipitate was formed**. Which ions formed the precipitate compound? Which solubility rules did you follow?